

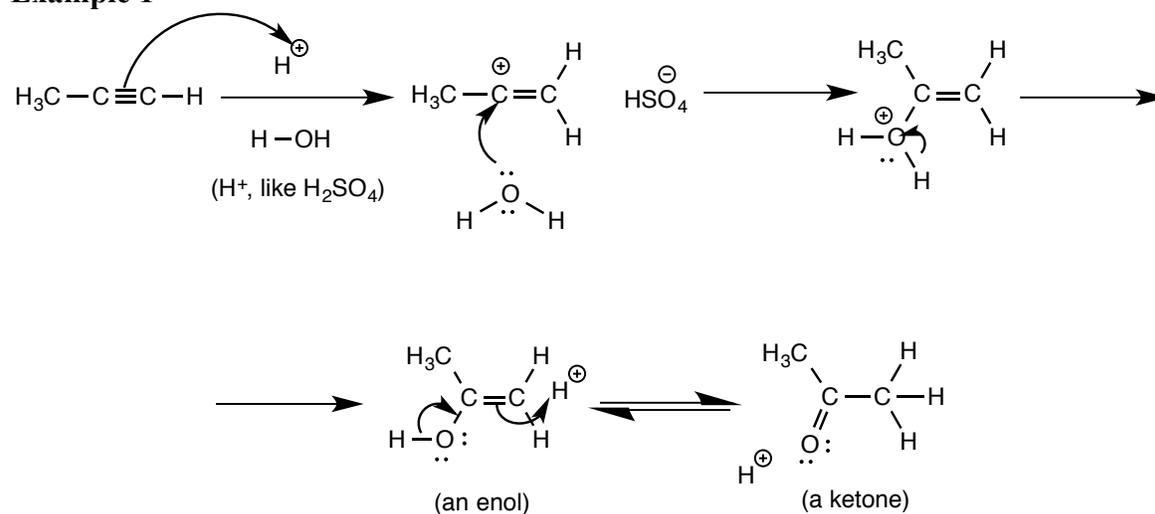
Upcoming Topics:

1. Alkyne Reactions
2. Radical Additions to Alkenes
3. Polymers
4. Conjugated Systems
5. Aromatics

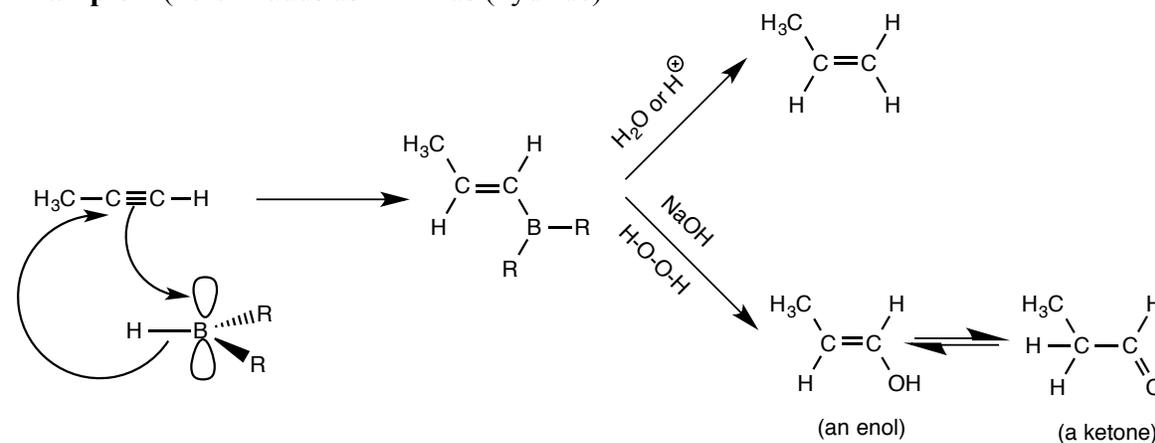
REVIEW

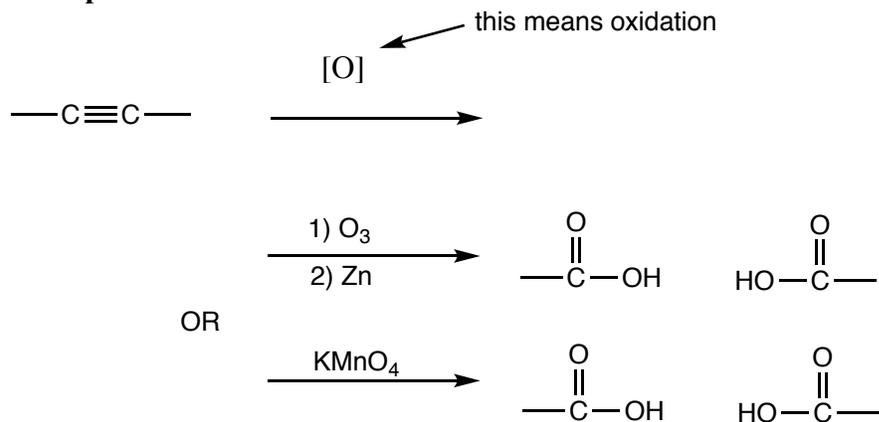
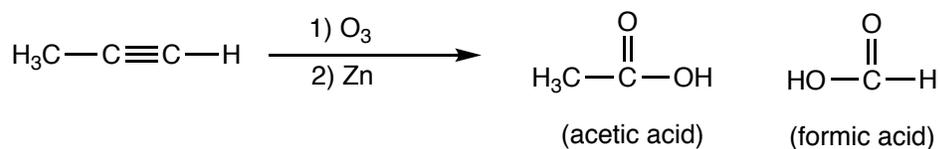
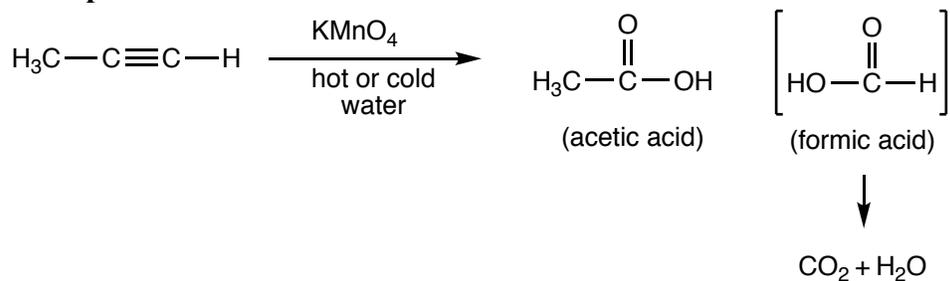
:Base picks up a proton (H^+) very fast, may be negatively charged

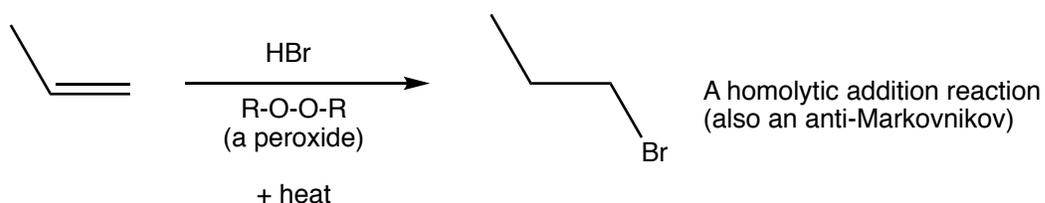
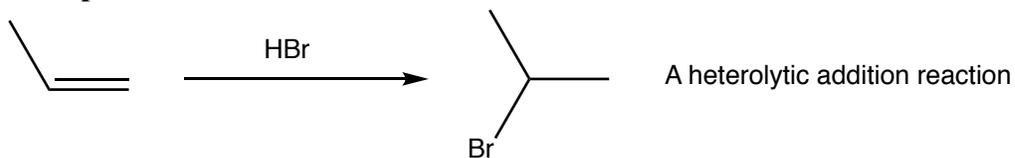
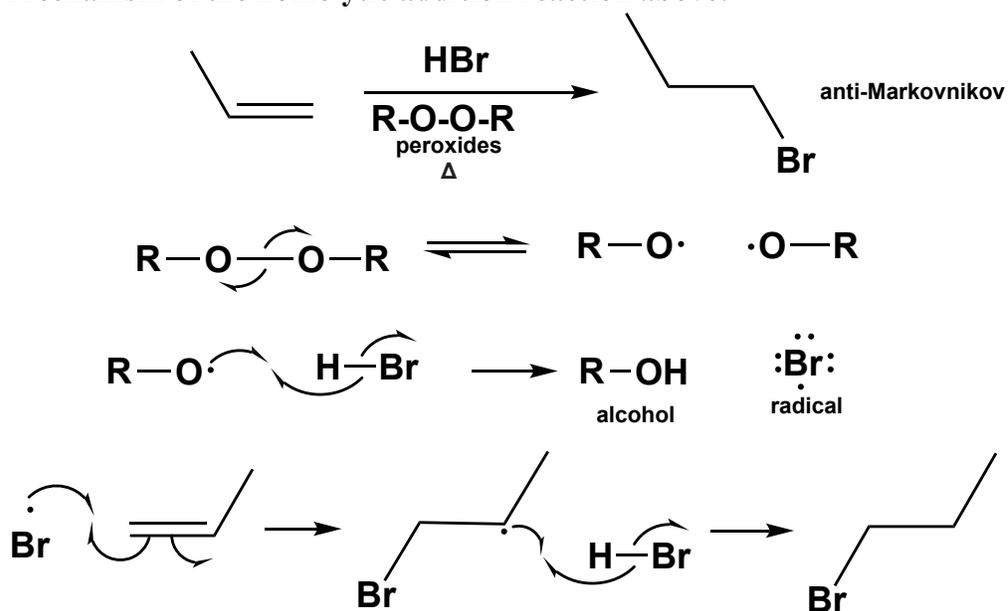
:Nucleophile attaches other atoms (like carbon)

Example 1

The Keto–Enol forms are called Tautomers (Rapidly interconverting of structural isomers)

Example 2 (here H adds as H minus (hydride))

Example 3**Example 4****Example 5****General Reaction Scheme**

Example 6**Mechanism of the homolytic addition reaction above:**

More highly substituted secondary (2°) radical formed

Polymers

poly – many, meros – parts

Natural Polymers:

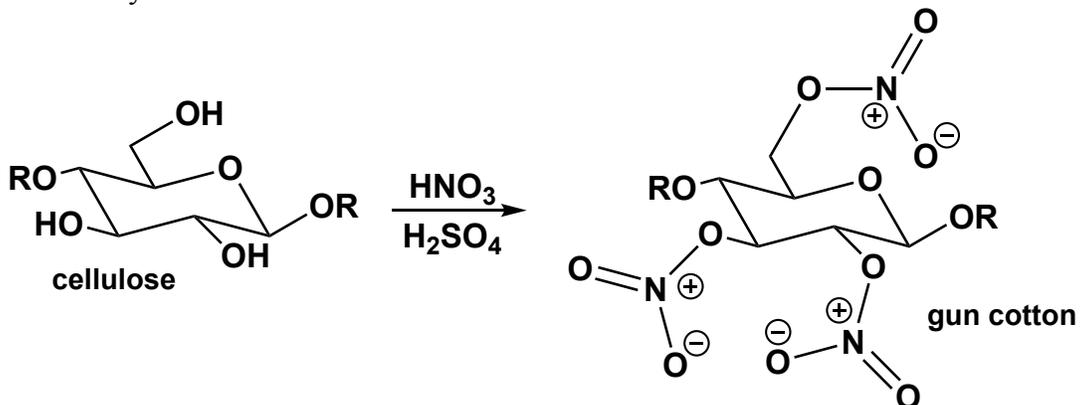
- Nucleic acids (DNA, RNA)
- Proteins and peptides (amino acid polymers)
- Polysaccharides (cellulose)
- Fats, polyketides (polymers of acetic acid)

Human-Made Polymers

- Christian Schönbein, 1826 professor in Basel
 - Mopped up a spill of H_2SO_4 and HNO_3 with an apron made of cotton (cellulose, a polymer of glucose),
 - He washed the apron and hung it up to dry
 - The apron exploded → He had created gun cotton!

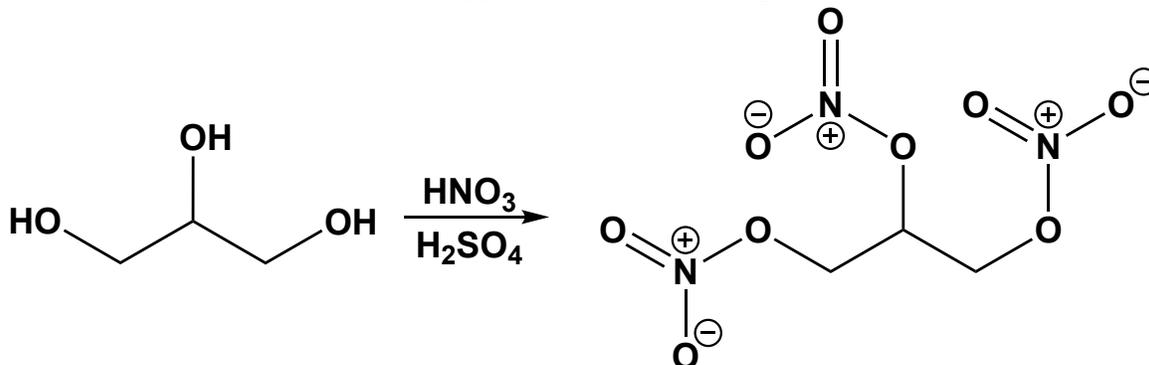
How did this happen?

In the presence of H_2SO_4 and HNO_3 , the free hydroxyl groups form a material that spontaneously combusts.



Another example: glycerol or glycerin

Under the same conditions, glycerol will form nitroglycerine

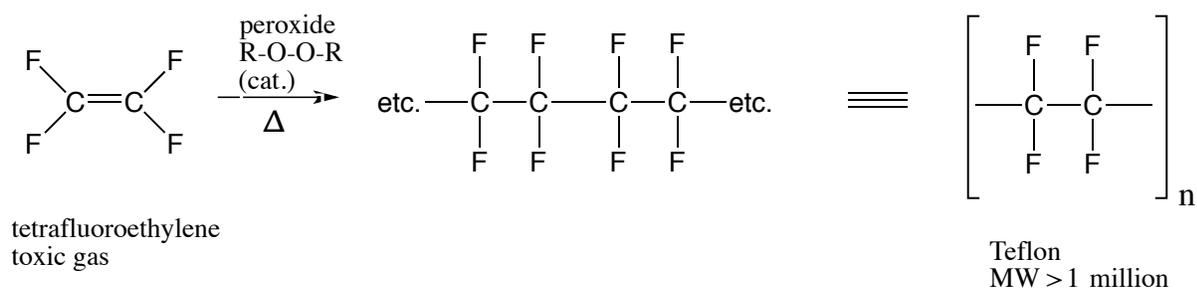


2 to 5 % of nitroglycerine was combined with diatomaceous earth → dynamite!

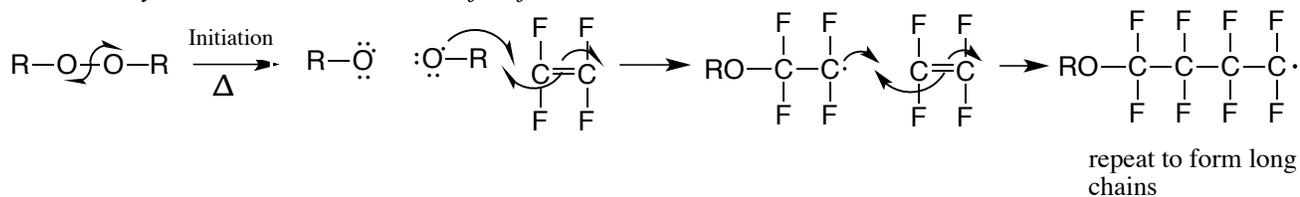
Who did this reaction? Alfred Nobel, the founder of the Nobel Prize!

- Wallace Carothers 1896-1937 DuPont
 - Nylon (polyamide)
 - Dacron (polyester)
- Roy Plunkett 1938 DuPont
 - Teflon (created from tetrafluoroethylene, a toxic gas)

Teflon: Polytetrafluoroethylene



Polymerization Mechanism of Teflon



PE: Polyethylene

