

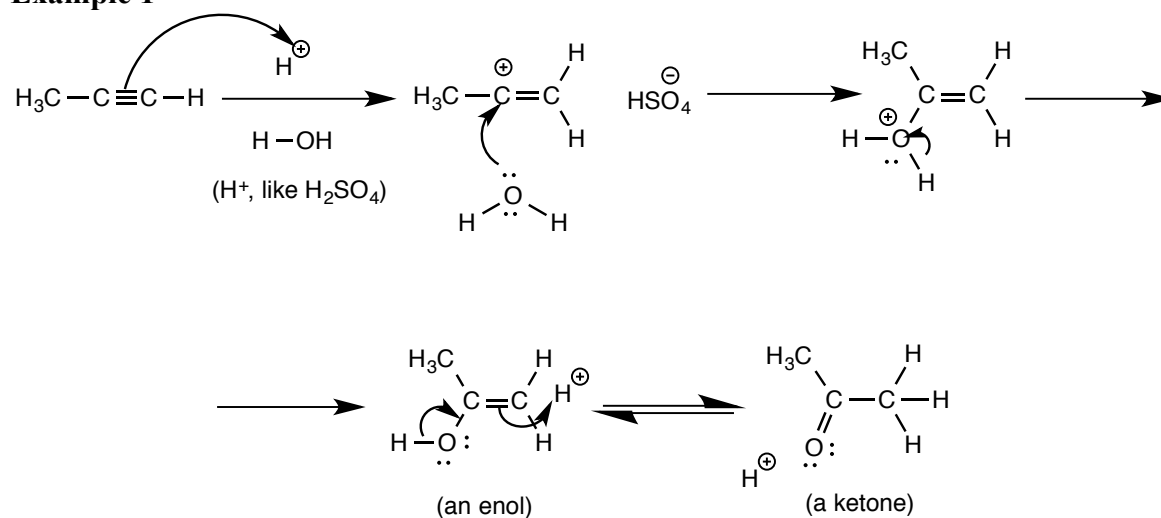
**Upcoming Topics:**

1. Alkyne Reactions
2. Radical Additions to Alkenes
3. Polymers
4. Conjugated Systems
5. Aromatics

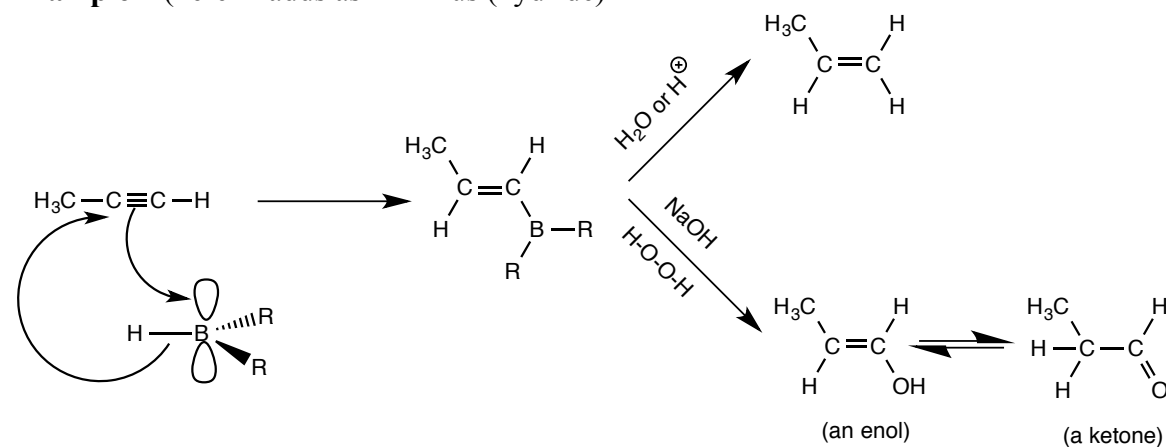
**REVIEW**

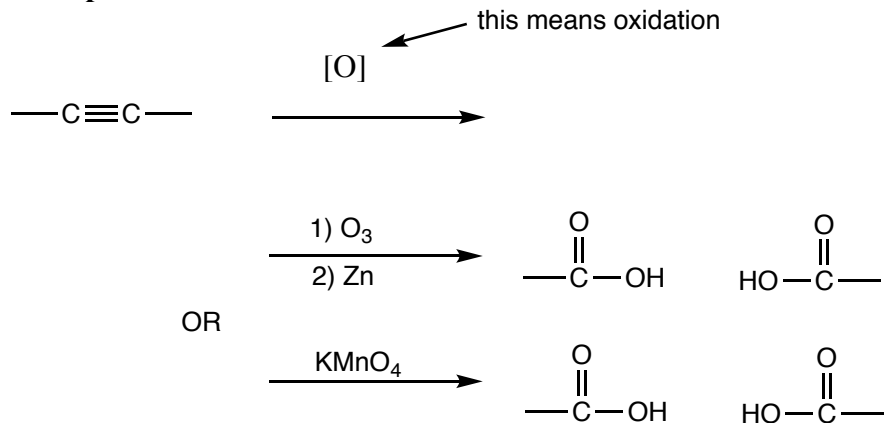
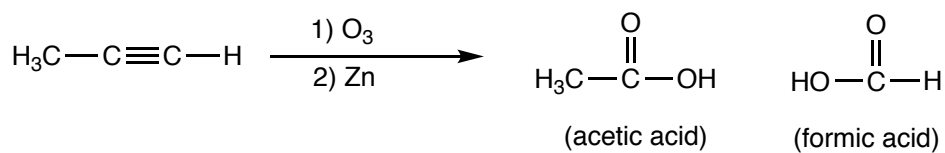
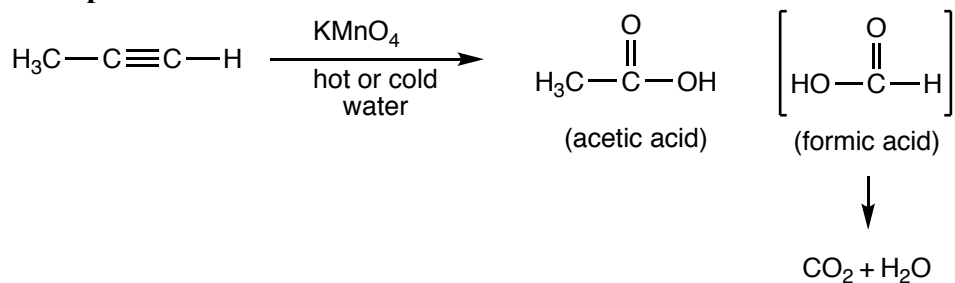
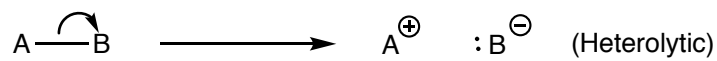
**:Base** picks up a proton ( $\text{H}^+$ ) very fast, may be negatively charged

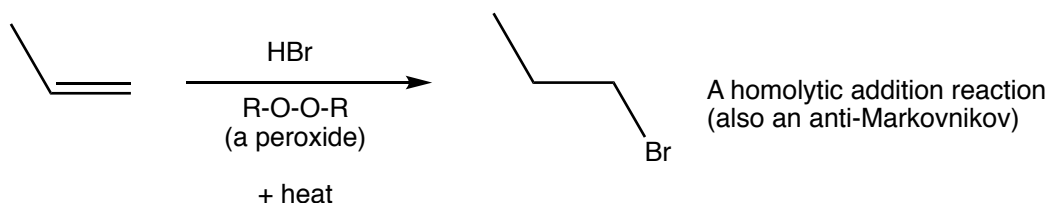
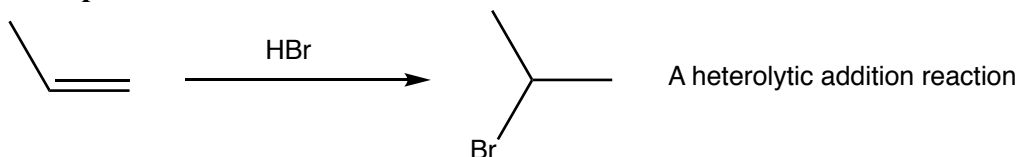
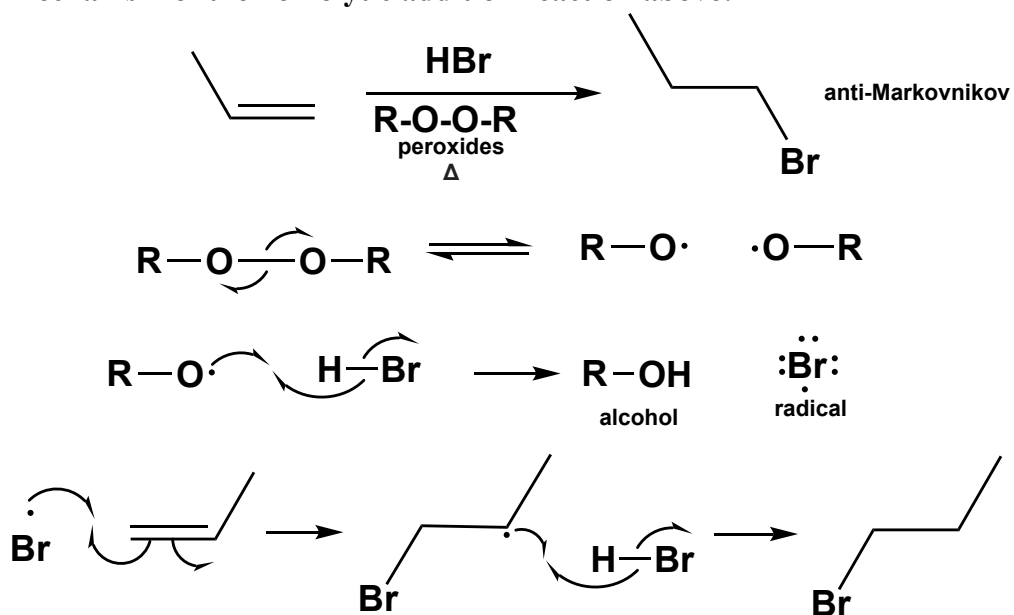
**:Nucleophile** attaches other atoms (like carbon)

**Example 1**

The Keto–Enol forms are called Tautomers (Rapidly interconverting of structural isomers)

**Example 2** (here H adds as H minus (hydride))

**Example 3****Example 4****Example 5****General Reaction Scheme**

**Example 6****Mechanism of the homolytic addition reaction above:****Polymers**

poly – many, meros – parts

*Natural Polymers:*

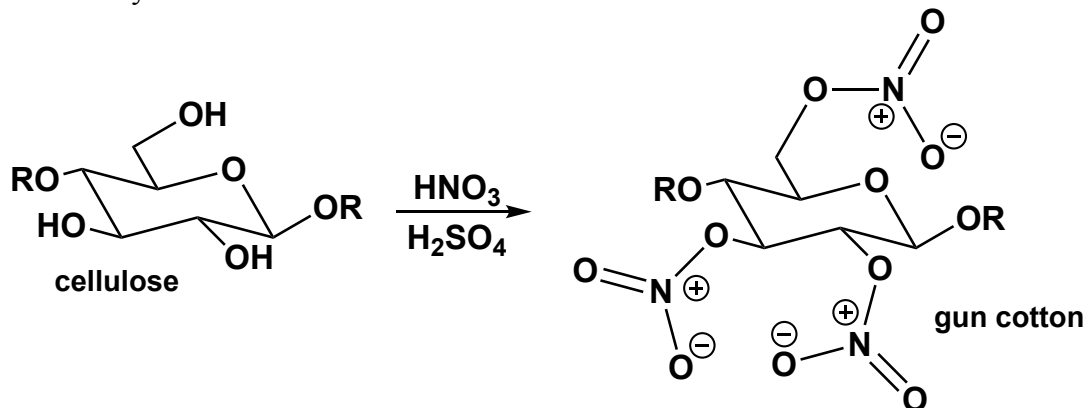
- Nucleic acids (DNA, RNA)
- Proteins and peptides (amino acid polymers)
- Polysaccharides (cellulose)
- Fats, polyketides (polymers of acetic acid)

### Human-Made Polymers

- Christian Schönbein, 1826 professor in Basel
  - Mopped up a spill of  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$  with an apron made of cotton (cellulose, a polymer of glucose),
  - He washed the apron and hung it up to dry
    - The apron exploded → He had created gun cotton!

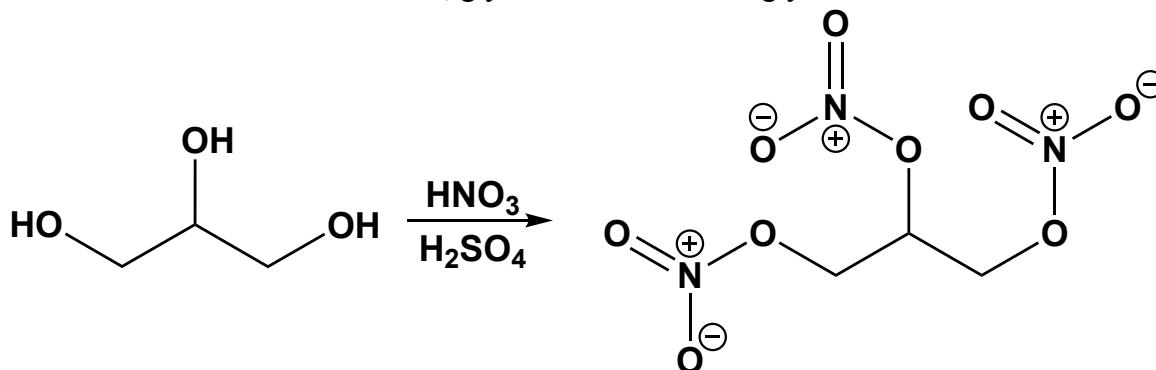
#### How did this happen?

In the presence of  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$ , the free hydroxyl groups form a material that spontaneously combusts.



#### Another example: glycerol or glycerin

Under the same conditions, glycerol will form nitroglycerine



2 to 5 % of nitroglycerine was combined with diatomaceous earth → dynamite!

*Who did this reaction?* Alfred Nobel, the founder of the Nobel Prize!

- Wallace Carothers 1896-1937 DuPont
  - Nylon (polyamide)
  - Dacron (polyester)
- Roy Plunkett 1938 DuPont
  - Teflon (created from tetrafluoroethylene, a toxic gas)

