

<b>Instructor:</b>	Dr. H. Taube								
<b>Office:</b>	Chem Bldg E5-33J								
<b>Phone:</b>	492-5239								
<b>Lecture Hours:</b>	T, R 12:30 - 13:50 h;								
<b>Office Hours:</b>	M, W, F 13:00 - 14:00 h; T, R 14:00 - 15:00 h; by appointment or take your chances and drop in								
<b>Email:</b>	<a href="mailto:htaube@ualberta.ca">htaube@ualberta.ca</a>								
<b>Course Mark Breakdown:</b>	<table><tr><td>Labs (pass mark 60%)</td><td>25 %</td></tr><tr><td>Quiz</td><td>11 %</td></tr><tr><td>MidTerm Exam</td><td>25%</td></tr><tr><td>Final Exam</td><td>39 %</td></tr></table> <p>A letter grade will be assigned based on your course average, using the attached approximate University Guidelines.</p>	Labs (pass mark 60%)	25 %	Quiz	11 %	MidTerm Exam	25%	Final Exam	39 %
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<b>Code of Student Behavior:</b>	No misrepresentation, plagiarism or cheating. See 2010/11 University Calendar, pp. 777 - 798, for details.								
<b>Text:</b>	"Organic Chemistry", Solomons & Fryhle, 9 <sup>th</sup> ed., 2008								
<b>Recommended Text:</b>	"Study Guide ....", Solomons & Fryhle, 9 <sup>th</sup> ed., 2008								
<b>Molecular Models:</b>	Extremely useful; "Darling" type recommended; available at Organic Chemistry Store, Chem W1-32								
<b>Lab Manuals:</b>	"Organic Chemistry Experiments" & "Laboratory Report Bookbook", Chem261, H. Wan, <b>fa 2010 ed.</b> , available at Organic Chemistry Store, Chem W1-32								
<b>Course Outline:</b>	Textbook chapters 1 - 8, 10, 11; selected sections; detailed checklists of topics will be provided during the course.								
<b>Course Schedule:</b>	See overleaf (includes exam dates).								
<b>Labs:</b>	Begin week of 13 Sep. 2010. <b>Safety glasses mandatory.</b> Read lab manual and prepare for the first lab during the week before. Room assignments to be posted by Fri, 10 Sep., near Chem E1-26.								
<b>Help Sessions:</b>	<b>Theory &amp; Lab:</b> Room E1-34, open hours: Monday 9 - 11, 1 - 5 Tuesday 9 - 5 Wednesday 9 - 5 Thursday 9 - 5 Friday 9 - 11, 1 - 3								
<b>Supplementary Information:</b>	• (lecture notes, homework, exam information, ....): <a href="http://www.chem.ualberta.ca/~htaube">www.chem.ualberta.ca/~htaube</a> • exam scores: on <i>eClass</i>								

Week		Theory		Labs
#	Beginning	Tuesday	Thursday	
01	06 Sep	—	Intro	—
02	13 Sep	Bonding	Bonding	Check-In/ Recrystall <sup>n</sup>
03	20 Sep	Bonding	Reaction Theory	Chromatography/ M.P.
04	27 Sep	FG's/ IR	FG's/ IR	Infrared Spectroscopy
05	04 Oct	Alkanes	Alkanes/ <b>Quiz</b>	Distillation
06	11 Oct	Conformation	Conformation	—
07	18 Oct	Stereochemistry	Stereochemistry	Trimyristin
08	25 Oct	Stereochemistry	Nucleophilic S & E	Analgesic Separations I
09	01 Nov	Nucleophilic S & E	<b>Midterm Exam</b>	Analgesic Separations II
10	08 Nov	Nucleophilic S & E	—	—
11	15 Nov	Nucleophilic S & E	Alkenes	Stereochemistry
12	22 Nov	Alkenes	Alkenes	Nucleophilic Subst <sup>n</sup>
13	29 Nov	Alkynes	Alcohols	Final Quiz/Check-out
14	06 Dec	Ethers	—.	—
<b>FINAL EXAM: Friday, 17 December 2010, 14:00 h</b>				
Deferred Exam (as required): Tuesday, 18 Jan 2011, 13:00 h, Room TBA				

# Preliminary/Approximate

## Guide for Grade Assignments

CHEM 164/261, A2

2010 09 07  
HT

Avg. Course Mark	Grade (num. equiv.)	# of Students
> 91 %	A+ (4.0)	5 %
> 86 %	A (4.0)	7 %
> 81 %	A- (3.7)	12 %
> 76 %	B+ (3.3)	15 %
> 71 %	B (3.0)	16 %
> 66 %	B- (2.7)	14 %
>61 %	C+ (2.3)	11 %
> 56 %	C (2.0)	8 %
>51%	C- (1.7)	5 %
>49%	D+ (1.3)	3 %
>45%	D (1.0)	2 %
<45%	F (0.0)	2 %

Abbreviation	Full Term	Abbreviation	Full Term
A/B	acid/ base	g	gaseous
acc	according to	gem	geminal
add <sup>l</sup>	additional	h, hi	high
amt	amt	i.e.	that is; (definition)
approx	approximate	ill	illustration
aq	aqueous	incl	including
atm	atmosphere	KE	kinetic energy
ax	axial	l	liquid
b/c	because	l, lo	low
B/L	Bronsted - Lowry	lge	large
c/w	complete with	lvl	level
cat	catalyst	mech	reaction mechanism
cmpd	compound	MF	molecular formula
coeff	coefficient	mvmt	movement
conc	concentrated	N.B.	nota bene; note carefully
config	configuration	neg	negative
conform	conformation	org	organic
conj	conjugated	ox <sup>n</sup>	oxidation
const	constant	PE	potential energy
def	definition	pos	positive
diff	different	re	regarding
dil	dilute	red <sup>n</sup>	reduction
DofU	degree of unsaturation	ref	reference
e <sup>-</sup>	electron	req <sup>d</sup>	required
e.g.	for example	res	resonance
e/n	electronegative	rev	reverse
Ex	example	r.t.	room temp., ~ 20°C
edg	electron donating group	rxn	reaction
emr	electromagnetic radiation	s	solid
eq	equatorial	sat <sup>d</sup>	saturated
equil	equilibrium	sol <sup>n</sup>	solution
equiv	equivalent	sub	substituent
equ <sup>n</sup>	equation	TBA	to be announced
ewg	e <sup>-</sup> withdrawing group	tot	total
exp	experimental	trtmt	treatment
FC	formal charge	vic	vicinal
FG	functional group	v.v.	vice versa; "opposite also true"
fig	figure	w/	with
fwd	forward	w/t	without
		xs	excess